

ppt Rxns

physical states

(s) - solid

(l) - liquid

(g) - gas

(aq) - aqueous (dissolved in water)

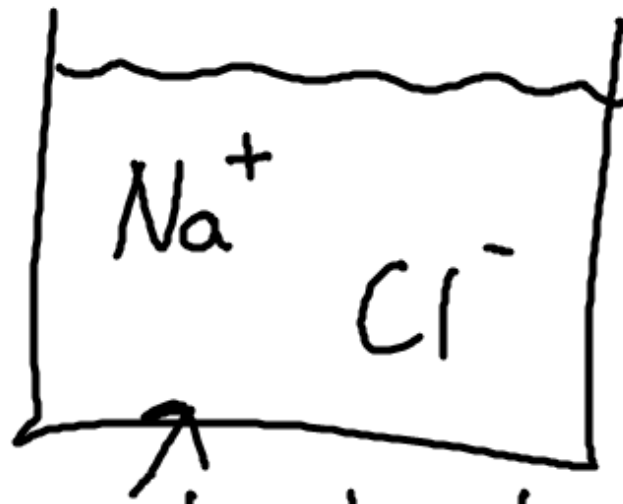
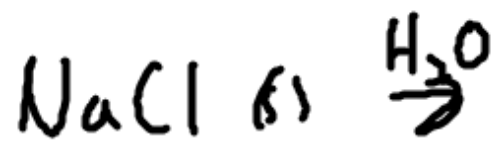
Solute - salt (gets dissolved)

Solvent - H_2O (does the dissolving)

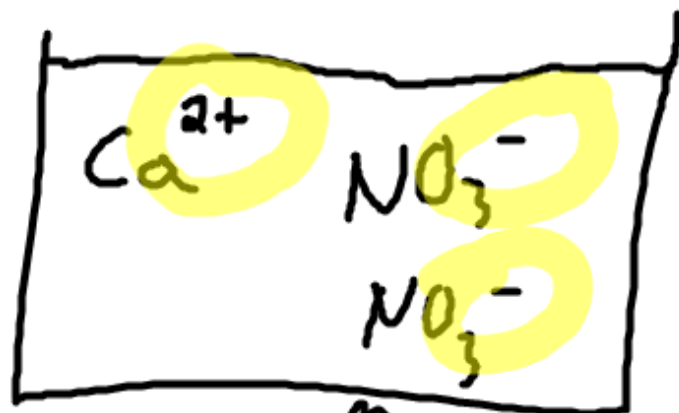
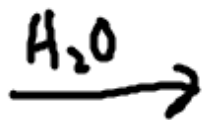
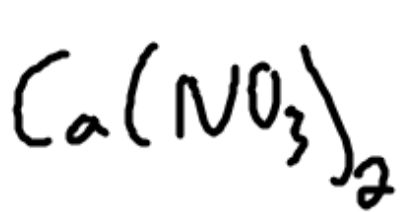
Soluble - can dissolve

insoluble - does not dissolve

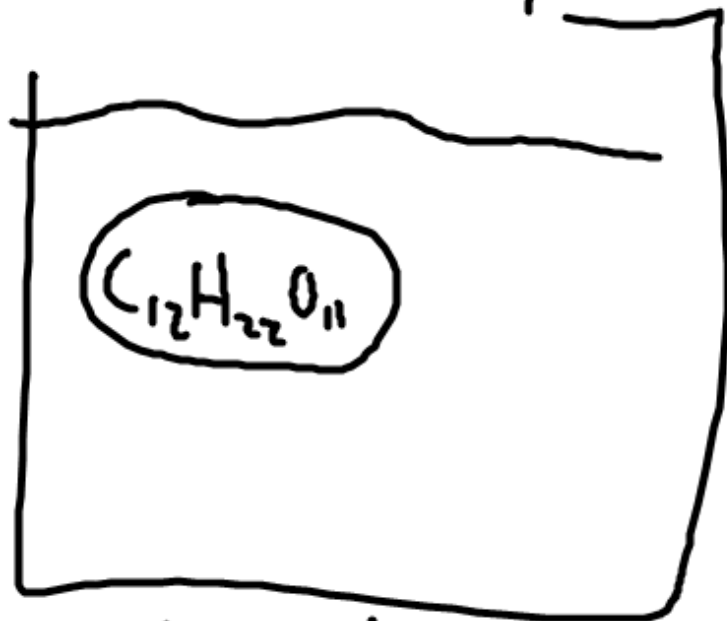
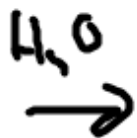
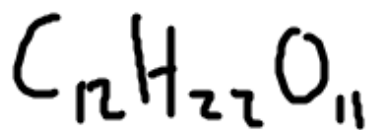
rock, Au,



dissociate - breaks up into ions



↗
electrolyte



nonelectrolyte

Not ionic

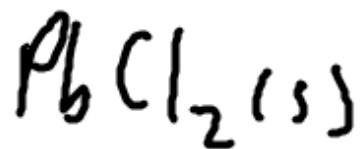
Solubility Rules

① NO_3^- always soluble

② Grp 1's + NH_4^+ soluble

③ most Halogens (Grp 17) soluble

except: Ag, Hg, Pb



④ most SO_4^{2-} soluble

except: Ba, Ca, Pb, Hg



most OH^- insoluble

expect -grp1 / NH_4^+

NaOH (aq)

most S^{2-} , CO_3^{2-} , PO_4^{3-} are insoluble

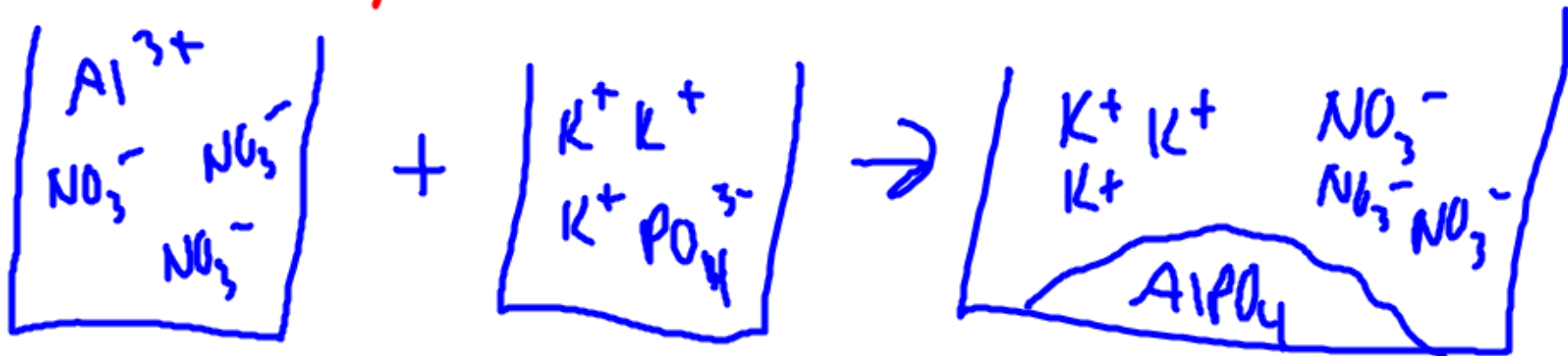
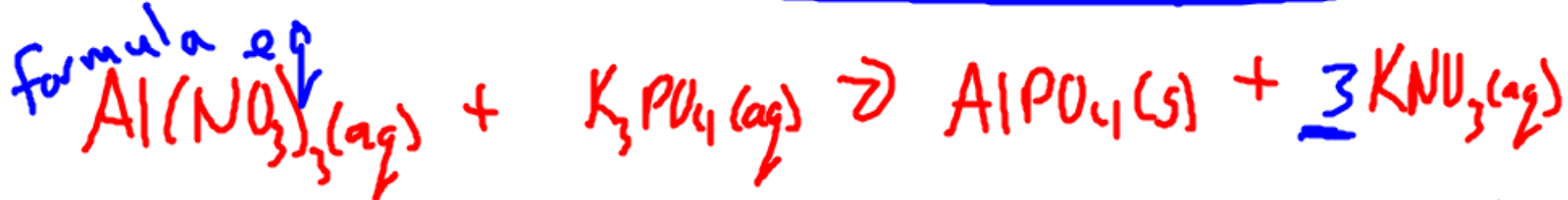
ppt rxn

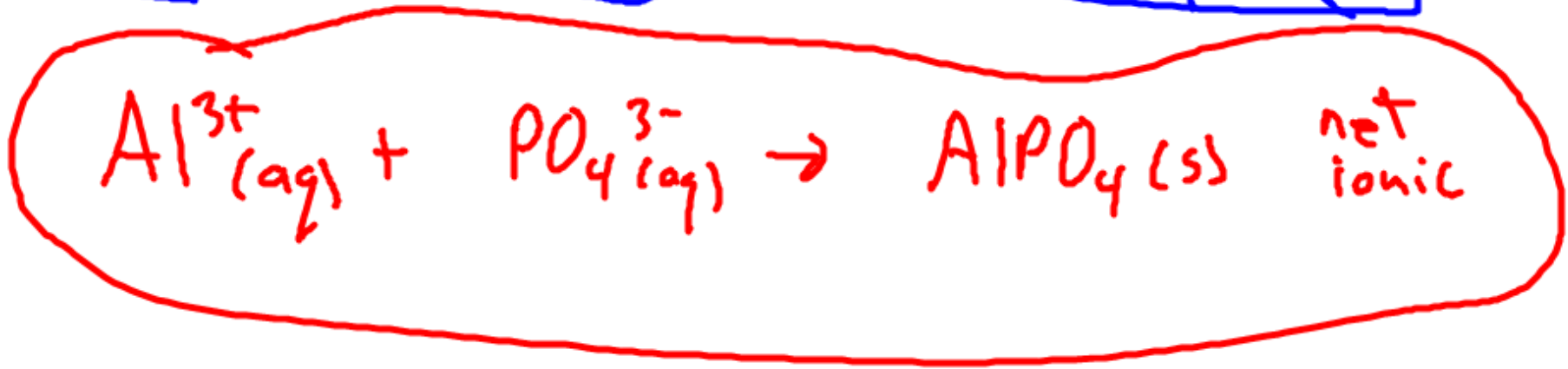
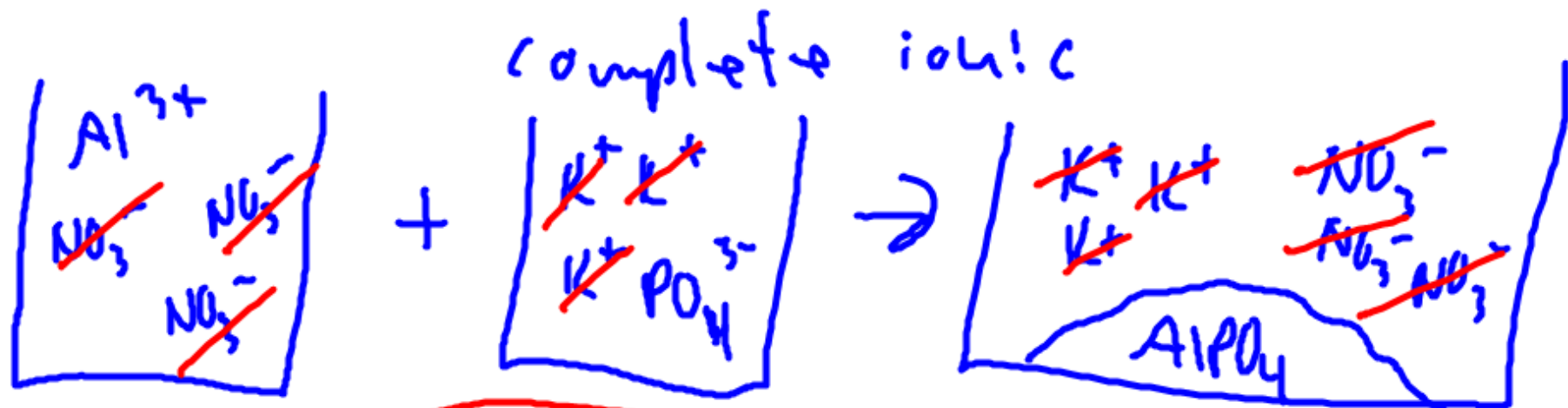
Solutions of Aluminium nitrate and potassium phosphate are mixed.

word eq

aluminium nitrate + potassium phosphate \rightarrow aluminium phosphate + potassium nitrate

formula eq





$NO_3^- + K^+$ spectator ions